

CHEMISTRY DATA SHEET

NAME	SYMBOL	VALUE
Standard pressure	p^θ	$1,013 \times 10^5 \text{ Pa}$
Molar gas volume at STP	V_m	$22,4 \text{ dm}^3 \cdot \text{mol}^{-1}$
Standard temperature	T^θ	273K
Charge on electron	e	$- 1,6 \times 10^{-19} \text{ C}$

$$n = \frac{m}{M}$$

$$n = \frac{N}{N_A}$$

$$c = \frac{n}{V} \quad \text{or} \quad c = \frac{m}{MV}$$

$$n = \frac{V}{V_m}$$

$$\frac{c_a V_a}{c_b V_b} = \frac{n_a}{n_b}$$

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1 \times 10^{-14} \quad \text{at } 298\text{K}$$

$$E_{\text{cell}}^\theta = E_{\text{cathode}}^\theta - E_{\text{anode}}^\theta$$

$$E_{\text{cell}}^\theta = E_{\text{reduction}}^\theta - E_{\text{oxidation}}^\theta$$

$$E_{\text{cell}}^\theta = E_{\text{oxidising agent}}^\theta - E_{\text{reducing agent}}^\theta$$